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Exercises 1. Determine the aluminum ion concentration and the chloride ion concentration in 0.125 M  $\text{AlCl}_3$ .

2. Determine the chromium ion

concentration and the sulfate ion

concentration in a 300 mL 0.050 g of

chromium(III) sulfate.

3. Determine the mass of metal ion in each of the

following: 500 mL of 0.752 M

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Cl<sub>2</sub> to produce 6.84 grams of a pure  
metal chloride. the metal chloride is  
dissolved in a 100.0-mL volumetric  
flask which is then filled up to the  
mark, the concentration of metal ions  
is

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$\text{BaSO}_4$  is 1 to 1 we can get grams

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